

Redox Reactions

NEET previous year questions · Chemistry

3 questions from the official papers: Re-NEET 2026 (2), NEET 2020 (1). Each is followed by its answer and a worked solution written by BrainIt. Generated on 23 October 2026.



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1. Re-NEET 2026 · Q57

In an acidic medium, 10 mL of 0.25 M oxalic acid is titrated with KMnO_4 solution. If the volume of KMnO_4 solution required to reach end point is 10 mL, the strength of the KMnO_4 solution is

- (1) 0.10 M (2) 0.20 M (3) 0.25 M (4) 0.15 M

Answer (1)

Sol. Oxalate is oxidised by two electrons ($\text{C}_2\text{O}_4^{2-} \rightarrow 2\text{CO}_2 + 2\text{e}^-$) so its n-factor is 2, while permanganate in ACIDIC medium is reduced by five electrons ($\text{MnO}_4^- \rightarrow \text{Mn}^{2+}$) so its n-factor is 5. Equating equivalents, $(0.25 \times 2)(10) = (M \times 5)(10)$, giving $M = 0.10$ M. NCERT Class 11, Redox Reactions (redox titrations).

2. Re-NEET 2026 · Q81

The correct decreasing order of oxidation state of the underlined atom in each molecule is

- (1) $\underline{\text{P}}\text{4O}_{10} > \underline{\text{S}}\text{O}_3 > \text{H}_2\underline{\text{O}}$
 (2) $\underline{\text{N}}\text{2O}_5 > \underline{\text{Al}}\text{2O}_3 > \text{H}_2\underline{\text{S}}$
 (3) $\underline{\text{Pb}}\text{O}_2 > \underline{\text{N}}\text{2O}_3 > \underline{\text{S}}\text{O}_3$
 (4) $\underline{\text{P}}\text{4O}_6 > \underline{\text{Cl}}\text{2O}_7 > \underline{\text{Al}}\text{H}_3$

Answer (2)

Sol. Assigning oxidation numbers in option (2): nitrogen in N_2O_5 is +5, aluminium in Al_2O_3 is +3, and sulphur in H_2S is -2 (hydrogen is +1 because sulphur is the more electronegative partner). The sequence +5 > +3 > -2 is genuinely decreasing, so (2) is correct. NCERT Class 11, Redox Reactions (oxidation number rules).

3. NEET 2020 · Q131

What is the change in oxidation number of carbon in the reaction $\text{CH}_4(\text{g}) + 4\text{Cl}_2(\text{g}) \rightarrow \text{CCl}_4(\text{l}) + 4\text{HCl}(\text{g})$?

- (1) 0 to +4 (2) -4 to +4
 (3) 0 to -4 (4) +4 to +4

Answer (2)

Sol. In CH_4 , the oxidation number of carbon is -4 (since hydrogen is +1). In CCl_4 , the oxidation number of carbon is +4 (since chlorine is -1). Thus, the change in oxidation number of carbon is from -4 to +4.

Chapter: Redox Reactions.

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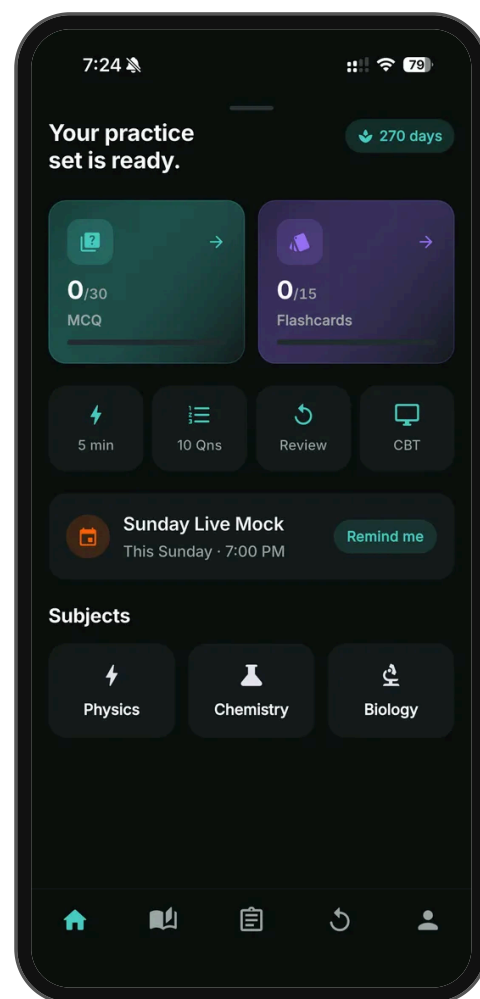
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